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Answers

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Questions pertaining to redox reactions If you're seeing this message, it means we're having trouble loading external resources on our website. If you're behind a web filter,

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**Redox reactions
questions (practice)
| Khan Academy**

Practice Problems:

Redox Reactions.

Determine the
oxidation number of
the elements in each of
the following

compounds: a. H_2CO_3

b. N_2 c. $\text{Zn}(\text{OH})_2$

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2-d. NO_2 2-e. LiH f. Fe^{3+}

O_4 Hint; Identify the

species being oxidized and reduced in each of the following reactions:

a. $\text{Cr} + \text{Sn}^{4+} \rightarrow \text{Cr}^{3+} + \text{Sn}^{2+}$

b. $3 \text{Hg}^{2+} + 2 \text{Fe}(\text{s}) \rightarrow 3 \text{Hg} + 2 \text{Fe}^{3+}$

c. $2 \text{As}(\text{s}) + 3 \text{Cl}_2(\text{g}) \rightarrow 2 \text{AsCl}_3$

Hint

Hint

Practice Problems:

Redox Reactions

Practice Problems:

Redox Reactions

(Answer Key)

Determine the

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oxidation number of
the elements in each of
the following

compounds: a. H_2CO_3 H: +1, O: -2, C: +4
b. N_2 N: 0
c. $\text{Zn}(\text{OH})_2$ Zn: 2+, H: +1, O: -2
d. NO_2 N: +3, O: -2
e. LiH Li: +1, H: -1
f. Fe_3O_4 Fe: +8/3, O: -2;

Identify the species
being oxidized and
reduced in each of the
following reactions:

**Practice Problems:
Redox Reactions**

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Answers

Here are a few practice problems to quiz yourself about redox reactions! Answers are below.

1. The potential energy in the products is greater than the reactants. What type of reaction is it? A) exergonic B) endergonic

2. The potential energy in the reactants is greater than the products. What type of reaction is it? A) exergonic B) endergonic

3.

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**RedOx_practice_prob
lems.pdf - Here are
a few practice ...**

Redox Reactions
Chapter Exam Take
this practice test to
check your existing
knowledge of the
course material. We'll
review your answers
and create a Test Prep
Plan for you based on
your results.

**Redox Reactions -
Practice Test**

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Answers

Questions & Chapter Exam ..

Problems 11-25. Only the examples and problems. Return to Redox menu. Problem #1: $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$.

Solution: 1) Balanced half-reactions: $6e^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$.
 $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$.

2) Equalize the electrons:

Balancing redox

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**reactions in acidic
solution: Problems**

#1-10

A redox reaction
always involves A. a
change in oxidation
number B. a change in
phase ... page 7 Redox
practice worksheet.

Problem-Attic format
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**Redox practice
worksheet**

Page 11/26

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Reaction Practice

Oxidation-Reduction

Balancing Additional

Practice Problems

Acidic Solution 1. $\text{Ag} +$

$\text{NO}_3^- \rightarrow \text{Ag}^+ + \text{NO}_2$.

$\text{Zn} + \text{NO}_3^- \rightarrow \text{Zn}^{2+} +$

NH_4^+ + 3. $\text{Cr}_2\text{O}_7^{2-} +$

$\text{C}_2\text{H}_4\text{O} \rightarrow \text{C}_2\text{H}_4\text{O}_2 +$

Cr^{3+} + 4. $\text{H}_3\text{PO}_2 + \text{Cr}_2\text{O}_7^{2-}$

$\rightarrow \text{H}_3\text{PO}_4 + \text{Cr}^{3+}$

Basic Solution

Oxidation-Reduction

Extra Practice

Practice exercises.

Balance the following

equations of redox

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Problems And
Answers

reactions: Assign oxidation numbers to all elements in the reaction. Separate the redox reaction into two half reactions. Balance the atoms in each half reaction. Add the two half-reactions together and cancel out common terms.

Balancing redox equations - Practice exercises

Redox reaction from dissolving zinc in

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copper sulfate.

Introduction to

galvanic/voltaic cells.

Electrodes and voltage
of Galvanic cell. ...

Practice:

Electrochemistry

questions. This is the
currently selected

item. Electrochemistry.

Redox reaction from
dissolving zinc in
copper sulfate.

**Electrochemistry
questions (practice)**

| Khan Academy

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When balancing redox reactions, the overall electronic charge must be balanced in addition to the usual molar ratios of the component reactants and products. This example problem illustrates how to use the half-reaction method to balance a redox reaction in a solution.

**Balance Redox
Reaction Example**

Page 15/26

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**Problem -
ThoughtCo**

13.2 Writing redox and half-reactions (ESCQY)
Redox reactions and half-reactions (ESCQZ)
Remember from Grade 11 that oxidation and reduction occur simultaneously in a redox reaction. The reactions taking place in electrochemical cells are redox reactions. Two questions should be asked to determine if a reaction is a redox

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Reaction Practice
reaction:

**Writing redox and
half-reactions |
Electrochemical ...**

Question: Tutored
Practice Problem
21.1.3 COOKIES
TOWARDS TRADE
Balance Redox
Reactions In Acidic
Solution. Close Problem
Balance The Reaction
Between H_2AsO_4 And
Culto Form HASO , And
Cu? In Acidic Solution.
When You Have

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Reaction Practice

Balanced The

Equation Using The

Smallest Integers

Possible, Enter The

Coefficients Of The

Species Shown.

**Solved: Tutored
Practice Problem**

21.1.3 COOKIES

TOWARDS TR ...

Practice Problems;

References; Oxidation-

Reduction or "redox"

reactions occur when

elements in a chemical

reaction gain or lose

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Problems And
Answers

electrons, causing an increase or decrease in oxidation numbers. The Half Equation Method is used to balance these reactions. In a redox reaction, one or more element becomes oxidized, and one or more element becomes reduced.

**Balancing Redox
Reactions: Examples
- Chemistry
LibreTexts**

Balancing Redox

Page 19/26

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Equations Method 1:

Oxidation number

method 1. Assign

oxidation numbers to

all elements in the

reaction 2. From the

changes in O.N.,

identify the oxidized

and reduced species 3.

Compute the number

of electrons lost in the

oxidation and gained in

the reduction from the

O.N. changes 4.

Multiply one or both of

these numbers by

appropriate

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Center**

Only the examples and problems Return to Redox menu. Points to remember: 1) Electrons NEVER appear in a correct, final answer. In order to get the electrons in each half-reaction equal, one or both of the balanced half-reactions will be multiplied by a factor. 2) Duplicate items are always removed. These

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Reaction Practice

items are usually the
electrons, water ..

Answers

Balancing redox reactions in basic solution

Practice Problems

Electrochemistry. 1.

What is the difference between an oxidation-reduction reaction and a half-reaction? 2.

What is the function of the salt bridge in an electrochemical cell? 3.

What is the criterion for spontaneous chemical

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change based on cell potentials? Explain. 4.

Problems And
Answers

CHM 112

Electrochemistry

Practice Problems

Balancing REDOX

Reactions: Learn and

Practice Reduction-

Oxidation reactions (or

REDOX reactions)

occur when the

chemical species

involved in the

reactions gain and lose

electrons. Oxidation

and reduction occur

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Reaction Practice

Problems and

Answers

simultaneously in order to conserve charge. We can “see” these changes if we assign oxidation numbers to the reactants and products.

Balancing REDOX Reactions: Learn and Practice

A redox reaction is a type of chemical reaction in which reduction and oxidation occur.

Because electrons are

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Problems and

Answers

transferred between chemical species, ions form. So, to balance a redox reaction requires not only balancing mass (number and type of atoms on each side of the equation) but also charge.

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