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8.4  $-5-17-22 = x10 = 2.05 \times 10$ . 17.  $4.1 \times 10$ . 4.1. b)  $5.4 \times 10$ .  
 $7.3 \times 10$ . 32.  $14 = 5.4$ .  $7.3 \cdot x10$ .  $32-14 = .74 \times 10$ .  $18 = 7.4 \times 10$ .

17. Information: Adding and Subtracting Using Scientific Notation. Whenever you add or subtract two numbers in scientific notation, you must make sure that they have the same exponents. Your answer will then have the ...

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3. Iodine has 17 more electrons than krypton and so you can

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begin by writing [Kr] followed by orbitals for 617 more electrons. After 4p comes 5s<sup>2</sup>... Iodine = [Kr]5s<sup>2</sup>4d<sup>10</sup>5p<sup>5</sup> Any of the noble gases can be used for abbreviations. Here are a few more examples: iron = [Ar]4s<sup>2</sup>3d<sup>6</sup> cesium = [Xe]6s<sup>1</sup> phosphorus = [Ne]3s<sup>2</sup>3p<sup>3</sup>

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Given your answer to number 7, you should be able to answer the following questions. a) Find the pH of a solution that has a pOH of 4.6.  $9.4$ ;  $\text{pH} + \text{pOH} = 14$   $\text{pH} = 14 - \text{pOH} = 9.4$  b) The  $[\text{H}^+]$

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O + ] of a solution is  $2.55 \times 10^{-3}$  M.

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7; since chlorine is in column 17 (or "7A") it has 7 valence

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electrons. 3. Since  $\text{CCl}_4$  is made up of one carbon and four chlorine atoms, how many total valence electrons does  $\text{CCl}_4$  have? Hint: add your answer to question 1 and four times your answer to question 2.  $4 + 4(7) = 4 + 28 = 32$  4. Verify that there are 32 electrons pictured in the ...

### **ChemQuest 23 : Drawing Covalent Compounds**

Download Ebook Chemquest 49 Answer Key Chemquest 49 Answer Key The Online Books Page: Maintained by the University of Pennsylvania, this page lists over one ... Iodine is 80%  $^{127}\text{I}$  17%  $^{126}\text{I}$ , and 3%  $^{128}\text{I}$ . Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that

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## **Grosse Pointe Public School System / GPPS Home**

24 ChemQuest 8 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_

Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

### **24 ChemQuest 8 Name:**

Calculate the average atomic masses. Round all answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80%  $^{127}\text{I}$  17%  $^{126}\text{I}$ , and 3%  $^{128}\text{I}$ . Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that have the following

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### **Chemquest 36 Answer Key - RTI**

118 ChemQuest 39 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_

Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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